

Writing Net Ionic Equations OR Brønsted-Lowry Equations for Acid-Base reactions5 Step Method

- List all species present
- Identify the SA and the SB
- Write an equation - transfer one proton
- Determine reaction type
- Determine the position of equilibrium

1. Write the Net Ionic Equation using Brønsted-Lowry theory for each aqueous reaction:

a) sodium carbonate and acetic acid

b) ammonia and nitrous acid

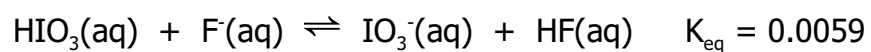
c) nitric acid and rubidium hydroxide

d) sulfuric acid and potassium phosphate

e) hydrofluoric acid and ammonium acetate

f) calcium chloride and lead sulfate

2. From the information below, identify the strongest acid and strongest base. Explain your choice.



3. Write the conjugate acid and/or base for the following:

Conjugate Acid						
Species	HCO_3^-	NO_3^-	H_3PO_4	H_2PO_4^-	S^{2-}	H_3AsO_4
Conjugate Base						