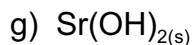
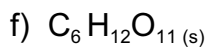
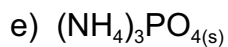
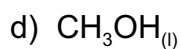
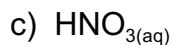
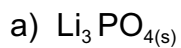


Dissociation and Ion Concentration

1. Write an equation to show what happens when each of the following dissolves in water.



2. Determine the concentration of aqueous sulfate ion in a 0.25 mol/L solution of $\text{Al}_2(\text{SO}_4)_{3(aq)}$.

3. What is the $[\text{OH}^-_{(aq)}]$ in 0.45 M strontium hydroxide?

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4. Determine the concentration of aqueous magnesium **and** aqueous acetate ions in a 0.75 M solution of magnesium acetate.

 5. The concentration of aqueous nitrate ion in a solution of $\text{Al}(\text{NO}_3)_3$ is determined to be 0.50 mol/L. What is the concentration of the aluminum nitrate solution?

 6. What is the concentration of aqueous ammonium ion in a solution prepared by dissolving 2.50 g of $(\text{NH}_4)_3\text{PO}_4$ in enough water to make 0.250 L of solution?

 7. What is the mass of $(\text{NH}_4)_2\text{CO}_3$ that must be dissolved in enough water to make a 500.0 mL solution in which the $[\text{NH}_4^+_{(\text{aq})}]$ is 0.19 mol/L?